## **Old Exam. Questions Ch.1**

# <u>T072:</u>

**Q1.:** The position y of a particle moving along the y axis depends on the time t according to the equation  $y = At - Bt^2$ . The dimensions of the quantities A and B are respectively: (Ans: L/T, L/T<sup>2</sup>)

## <u>T071</u>:

Q1.: A swimming pool is filled with 16,500 ft<sup>3</sup> of water. What is the volume of water in m<sup>3</sup>? (12 inch = 1 ft and 2.54 cm = 1 inch). (Ans: 467 m<sup>3</sup>)

**Q2.:** The position x of a particle is given by  $x = B t^2 + \frac{C}{B} t$ , where x is in meters and t is in seconds. The dimension of C is: (Ans:  $\frac{L^2}{T^3}$ )

# <u>T062:</u>

Q1: From the fact that the average density of the Earth is 5.50  $g/cm^3$  and its mean radius is  $6.37 \times 10^6 m$ , the mass of the Earth is: (Ans:  $5.95 \times 10^{24} kg$ )

**Q2:** Suppose  $A=B^n/C^m$  where *A* has dimensions LT, *B* has dimensions  $L^2T^{-1}$ , and *C* has dimensions  $LT^2$ . Then the exponents *n* and *m* have the values: (Ans: n = 1/5; m = -3/5)

# <u>T061</u>

Q1: An aluminum cylinder of density 2.70 g/cm<sup>3</sup>, a radius of 2.30 cm, and a height of 1.40 m has the mass of: (Ans: 6.28 kg)

## <u>T052</u>:

**Q1**. A nucleus of volume 3.4 x 10<sup>3</sup> fm <sup>3</sup> and mass of 1.0x 10<sup>2</sup> u has a density of: (1 fm =  $10^{-15}$  m, 1 u = 1.7 x  $10^{-27}$  kg) (**Ans:**  $5.0 \times 10^{16}$  kg/m <sup>3</sup>)

# <u>T051</u>:

**Q1**. The mass of 1.0 cm<sup>3</sup> of gold is 19.3 g. What is the mass of a solid cube of gold having a side of 0.70 cm ? (**Ans:**  $6.6 \times 10^{-3}$  kg)

#### <u>T042</u>:

 $\overline{\mathbf{Q1}}$  Express speed of sound, 330 m/s in miles/h . (1 mile = 1609 m )( **Ans:** A1 738 miles/h )

**Q2** A cylindrical can, 6.00 inches high and 3.00 inches in diameter is filled with water. Density of water is  $1.00 \text{ g/cm}^3$ . What is the mass of water in the can in gram? (1 inch = 2.54 cm. (**Ans:** 695 g.)

## <u>T041</u>:

 $\overline{\mathbf{Q1} \ 1}$  shake =  $10^{-8}$  seconds. Find out how many nano seconds (ns) are there in 1 shake.(1 nano =  $10^{-9}$ ) (**Ans:** 10 ns)

**Q2** A drop of oil (mass = 0.90 milligram and density = 918 kg/m<sup>3</sup>) spreads out on a surface and forms a circular thin film of radius = 41.8 cm and thickness h (see Fig 8). Find h in nano meter (nm). (1 nano =  $10^{-9}$ ) (**Ans:**1.8 nm)

#### <u>T032</u>:

**Q1** A solid lead cylinder has a mass of 56.5 kg and radius of 35 cm. Find the height of the cylinder. (The density of lead is  $11.3 \text{ g/cm}^3$ ) (**Ans:** 1.3 cm)

## <u>T031:</u>

 $\overline{\mathbf{Q1}}$  An empty fuel tank of a car needs 50 liters of gasoline to fill up. Find the volume of the fuel tank in m<sup>3</sup>. (1 milliliter = 1 cm<sup>3</sup>) (**Ans:** 0.050)

## **T022:**

**Q4** Dimension of an atom is often measured in a unit called Angstrom which is equal to 0.1 nm. 1 mm is equal to:  $(1 \text{ nm} = 10^{-9} \text{ m})$  (**Ans:**10 000 000 Angstrom )

**Q5** A student remembers that it takes roughly 8.4 minutes for the sun's light to reach the earth. Using this information and the fact that the speed of light is  $(3.0 \times 10^8)$  m/s, estimate the distance to the sun in km (**Ans:**  $1.50 \times 10^8$  km)

#### **T021:**

**Q1** The standard kilogram is a platinum-iridium cylinder 39 mm in height and 19.5 mm in radius. What is the density of the material? (Ans:  $21 \text{ g/cm}^3$ )

**Q3** The speed of sound in air is about 350 m/s. Express this speed in miles per hour (mi/h). (1 mile = 1.61 km) (**Ans:** 783 mi/h)

## <u>T012:</u>

**Q1** Speed of sound is 340 m/s. Express this in millimeters per nanosecond[ 1 ns =  $10^{-9}$  s]. (**Ans:** 3.40 x  $10^{-4}$  mm/ns)

## **T011:**

**Q1** Speed of sound is 330 m/s. Express this in miles per hour (1 mile = 1609 m). (Ans: 738 miles/h)

**Q2** The average radius of a nucleus is R = 10.0 fm. Find the density of the nucleus which has a mass of 15u [1 fm =  $10^{-15}$  m, 1 u = 1.66 x  $10^{-27}$  kg]. (Ans: 5.94 x  $10^{15}$  kg/m<sup>3</sup>)

## <u>**T992:**</u>

**Q1** A cube of copper has a mass m = 126 g. Find the number of copper atoms in this cube. Atomic mass of copper = 63.0 g/mole; Avogadro number =  $6.02 \times 10^{23}$  atoms/mole (**Ans:** 1.20 x 10<sup>24</sup>)

## <u>T991:</u>

 $\overline{\mathbf{Q2}}$  How many molecules of water are there in a cup containing 250 cm<sup>3</sup> of water? Molecular mass of H<sub>2</sub>O = 18 g/mole; Density of water = 1.0 g/cm<sup>3</sup>; Avogadro s number = 6.02 x 10<sup>23</sup> molecules/mole (**Ans:** 8.4 x 10<sup>24</sup>)

**Q3** Using the fact that the speed of light in space is about  $3.00 \times 10^8$  m/s, determine how many miles light will travel in one hour. (1 mile = 1.61 km) (**Ans:** 6.71 x 10<sup>8</sup> miles)