

PYP 001
Quiz # 6

Name: _____ ID# _____

Fill in the blanks:

1. Find the atomic mass of element X if it has three isotopes: X-30 with abundance of 60%, and mass 29.521 amu, isotope X-32 with abundance 29% and mass 28.572 amu and isotope X-35 with abundance 11% and mass 34.225 amu.

Solution:

$$29.521 \times \frac{60}{100} + 28.572 \times \frac{29}{100} + 34.225 \times \frac{11}{100} = 29.76323 \text{ amu}$$

2. Group number 2 elements belong to the alkaline-earth metals group.
3. Elements that mix with water to produce slippery solution and fire resistant are called alkaline-earth metals
4. The following elements in period 4 which one has the largest size
a. ${}_{36}\text{Kr}$ b. ${}_{30}\text{Zn}$ c. ${}_{19}\text{K}$ d. ${}_{20}\text{Ca}$

5. complete the following table

Isotope	Number of		
	Electrons	Protons	neutron
${}_{30}^{62}\text{Zn}$	30	30	32
${}_{55}^{136}\text{Cs}$	55	55	81
${}_{36}^{79}\text{Kr}$	36	36	43

