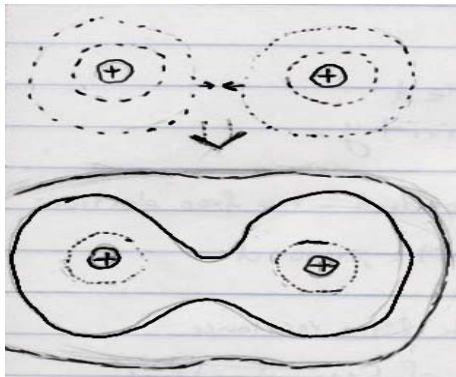


## Semiconductors

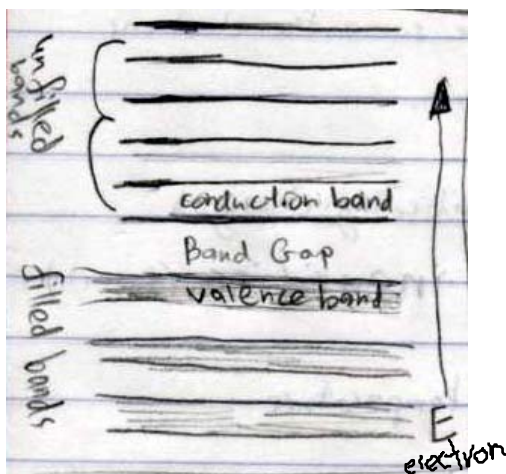
- Solid materials are classified in terms of electric conductivity.
  1. Insulators: do not conduct – no free electrons. e.g. Plastic, Wood.
  2. Conductors: conduct with little resistance – abundance of free electrons. e.g. Metals.
  3. Semiconductors: conduct but with large resistance – have some free electrons. e.g. Silicon.
- Resistivity can be significantly changed by adding certain impurities ( Also called Dopants).
- Resistivity depends on temperature.

## Energy Bands

When a solid is formed, energy levels in the atoms interact and form energy bands.

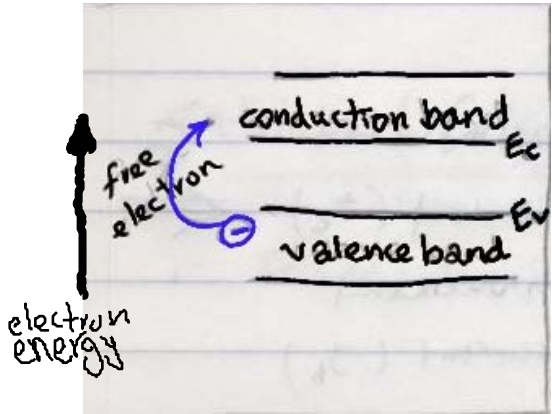


Energy Levels Splitting to form Bands



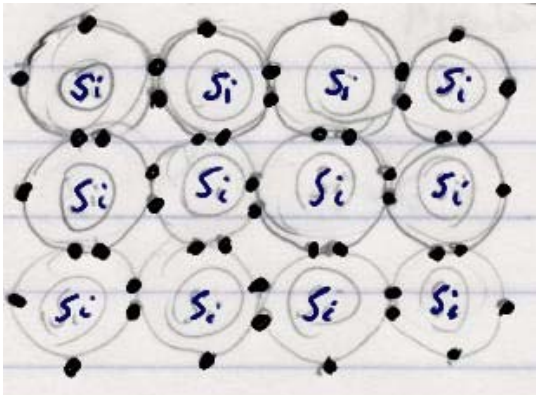
The Energy Bands

- the most important bands are:
  1. the valence band:
    - This is the outermost band filled with electrons. This is also the band where electrons that are involved in chemical bonds exist
  2. the conduction band:
    - this is the band next to the valence band where free electrons would be if they exist. It is the lowest of the unfilled bands
- there is an energy gap between these two bands which determine whether the solid is ( conductor / insulator / semiconductor )
- *Insulators* have large energy Gap ( $E.G > 2 \text{ e.v}$ )
- *Semiconductors* have smaller energy Gap ( $0 < E.G < 2 \text{ e.v}$ )
- *Conductors* have the valence & conduction bands overlapped ( $E.G < 0 \text{ e.v}$ )

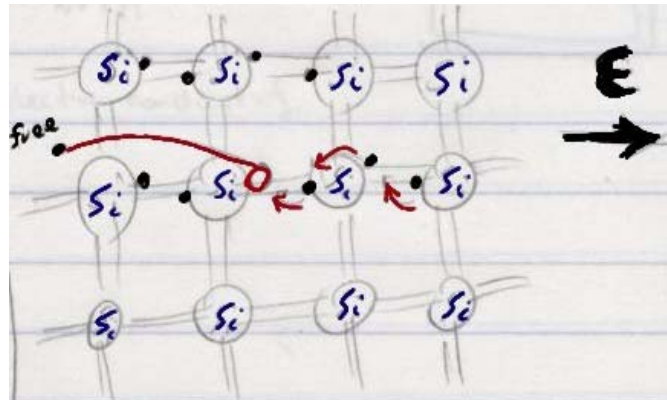


- The highest energy level of the valence band is called  $E_v$ .
- The lowest energy level of the conduction band is called  $E_c$ .
- $E_G = E_c - E_v$
- Energy Gap is a unique characteristic of a material, However, it can be slightly modified by applying physical pressure.

- An electron can become free by acquiring an energy ( thermal / photons ) =  $E_G$ 
  - It jumps from valence band to conduction band leaving a hole in the valence band.
  - This hole is positively charged & it is free to move within the valence band.
  - In Semiconductors, currents are transported via:
    1. electrons ( -ve charges)
    2. holes ( +ve charges)



Crystal structure, showing valence electrons associated with each bond



movement of a hole

