

CHAPTER 2

Force and Weight: $F = ma/g_c$, $W = mg/g_c$

Standard Acceleration of Gravity:

$$g = 9.80665 \text{ m/s}^2 = 980.665 \text{ cm/s}^2 = 32.174 \text{ ft/s}^2$$

$$g_c \text{ (gravitational conversion factor)} = 1 \text{ kg}\cdot\text{m}\cdot\text{s}^{-2}/\text{N} = 32.174 \text{ lb}_m\cdot\text{ft}\cdot\text{s}^{-2}/\text{lb}_f = 980.665 \text{ g}_m\cdot\text{cm}\cdot\text{s}^{-2}/\text{g}_f$$

Two Point Linear Interpolation:

$$y = y_1 + \frac{x - x_1}{x_2 - x_1}(y_2 - y_1)$$

Fitting a Straight line: $y = ax+b$, SLOPE, $a = \frac{y_2 - y_1}{x_2 - x_1}$, INTERCEPT, $b \begin{cases} = y_1 - ax_1 \\ = y_2 - ax_2 \end{cases}$

Method of Least Squares:

$$s_x = \frac{1}{n} \sum_{i=1}^n x_i \quad s_{xx} = \frac{1}{n} \sum_{i=1}^n x_i^2$$

$$s_y = \frac{1}{n} \sum_{i=1}^n y_i \quad s_{xy} = \frac{1}{n} \sum_{i=1}^n x_i y_i$$

$$\text{SLOPE} = \frac{s_{xy} - s_x s_y}{s_{xx} - (s_x)^2}, \quad \text{INTERCEPT} = \frac{s_{xx} s_y - s_{xy} s_x}{s_{xx} - (s_x)^2}$$

CHAPTER 3

$$\text{SG} = \rho/\rho_{\text{ref}}$$

$$\rho_{\text{ref}} (\text{H}_2\text{O}, 4^\circ\text{C}) = 1.000 \text{ g/cm}^3 = 1000 \text{ kg/m}^3 = 62.43 \text{ lb}_m/\text{ft}^3$$

Mass Fraction: $x_i = \frac{\text{mass of A}}{\text{total mass}}$, Mole Fraction: $y_i = \frac{\text{mole of A}}{\text{total mole}}$

Average Molecular Weight: $\bar{M} = \frac{m_i}{n_i} = \frac{\text{total mass of sample}}{\text{total moles in sample}}$

$$\bar{M} = y_1 M_1 + y_2 M_2 + \dots = \sum_{\text{all components}} y_i M_i$$

$$\frac{1}{\bar{M}} = \frac{x_1}{M_1} + \frac{x_2}{M_2} + \dots = \sum_{\text{all components}} \frac{x_i}{M_i}$$

Pressure: Hydrostatic Pressure: $P = P_o + \rho \frac{g}{g_c} h$

Head of Fluid: $P_h (\text{head of fluid}) = h_{\text{head of fluid}}$

$$P_{\text{absolute}} = P_{\text{gauge}} + P_{\text{atmospheric}}$$

General Manometer Equation: $P_1 + \rho_1 \frac{g}{g_c} d_1 = P_2 + \rho_2 \frac{g}{g_c} d_2 + \rho_f \frac{g}{g_c} h$

Differential Manometer Equation.: $P_1 - P_2 = (\rho_f - \rho) \frac{g}{g_c} h$

Manometer Formula for Gases: $P_1 - P_2 = h$

Temperature: $T(\text{K}) = T(^{\circ}\text{C}) + 273.15$, $T(^{\circ}\text{R}) = T(^{\circ}\text{F}) + 459.67$, $T(^{\circ}\text{R}) = 1.8T(\text{K})$, $T(^{\circ}\text{F}) = 1.8T(^{\circ}\text{C}) + 32$
 $1 \Delta\text{K} = 1 \Delta^{\circ}\text{C} = 1.8 \Delta^{\circ}\text{R} = 1.8 \Delta^{\circ}\text{F}$